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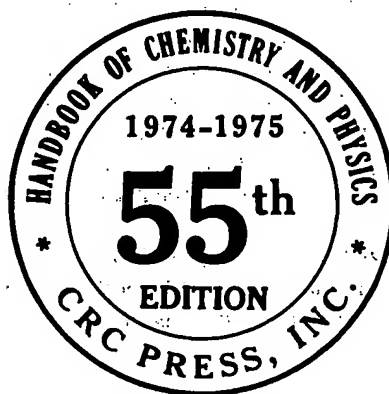
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EXHIBIT I



# Handbook OF Chemistry and Physics

A Ready-Reference Book of Chemical and Physical Data



EDITOR

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In collaboration with a large number of professional chemists and physicists whose assistance is acknowledged in the list of general collaborators and in connection with the particular tables or sections involved.

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# PHYSICAL CONSTANTS OF INORGANIC COMPOUNDS (Continued)

No.	Name	Synonyms and Formulae	Mol. Wt.	Crystalline form, properties and index of refraction	Density or specific gravity	Melting point, °C	Solubility, in grams per 100 cc			
							Boiling point, °C	Cold water	Hot water	Other solvents
Sodium										
a289	ferrate (III)	Ferrite, $\text{Na}_2\text{FeO}_4$	221.67	br, hex pl. or need	4.05	—	—	—	—	—
a290	ferricyanide	$\text{Na}_3\text{Fe}(\text{CN})_6 \cdot 3\text{H}_2\text{O}$	298.92	red cr, deliq	—	—	—	—	—	—
a291	ferrocyanate	Yellow prussiate of soda, $\text{Na}_4\text{Fe}(\text{CN})_6 \cdot 10\text{H}_2\text{O}$	484.04	pa yel, monocl, 1.518, 1.530, 1.544	1.458	—	—	—	—	—
a292	fluoaluminate	$\text{Na}_3\text{AlF}_6$	209.94	col, monocl, 1.304	2.90	1000	—	—	—	—
a293	fluoantimonate	$\text{Na}_3\text{SbF}_6$	258.73	rhomb, 3.375	<1360	—	—	—	—	—
a294	fluoberyllate	$\text{Na}_2\text{BeF}_4$	130.99	wh, rhomb or monoc	2.47	384	—	—	—	—
a295	fluoborate	$\text{NaBF}_4$	109.79	wh, rhomb, 2.47	384	—	—	—	—	—
a296	fluoride	Nat. fluorite, $\text{NaF}$	58.10	col, cub or tet, 1.336	2.558	—	1095	—	—	—
a297	fluoride, hydrogen	$\text{NaF} \cdot \text{H}_2\text{SO}_4$	100.99	col, or wh crs, powd, rhdr	2.08	—	—	—	—	—
a298	fluoride, orthophosphate	$\text{NaF} \cdot \text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}$	422.11	—	2.2165	—	—	—	—	—
a299	fluoroacetate, mono-	$\text{NaC}_2\text{H}_3\text{FO}_2$	100.02	wh powd, 1.02	—	200	—	—	—	—
a300	fluorophosphate, hexa-	$\text{Na}_6\text{P}_2\text{F}_{10}$	185.97	—	2.369	—	—	—	—	—
a301	fluorophosphate, mono-	$\text{Na}_2\text{PO}_3\text{F}$	143.95	col, 1.309	—	625	—	—	—	—
a302	fluosilicate	$\text{Na}_2\text{SiF}_6$	188.06	col, hex, 1312	2.679	—	—	—	—	—
a303	fluosulfonate	$\text{NaSO}_3\text{F}$	122.05	shiny leaf, hydr	—	red heat	—	—	—	—
a304	formaldehyde-sulfoxylate	$\text{NaHSO}_3 \cdot \text{CH}_2\text{O} \cdot 2\text{H}_2\text{O}$	154.12	rhomb pr, hydr	—	—	—	—	—	—
a305	formate	$\text{NaCHO}_2$	68.01	col, monocl, deliq	1.92	253	—	—	—	—
a306	2-furanacrylate	$\text{NaC}_4\text{H}_3\text{O}_5$	160.10	lt brn powd.	1.919	—	—	—	—	—
a307	metagermanate	$\text{Na}_2\text{GeO}_3$	166.57	wh, monocl, deliq, 1.59	3.31	1083	—	—	—	—
a308	metagermanate, heptahydrate	$\text{Na}_2\text{GeO}_3 \cdot 7\text{H}_2\text{O}$	292.68	col, rhomb, 1.309	—	83	—	—	—	—
a309	(mono-) $\alpha$ -glutamate	$\text{NaC}_5\text{H}_7\text{NO}_4$	169.11	wh cr, 1.309	—	red heat	—	—	—	—
a310	glycerophosphate, monohydrate	$\text{Na}_2\text{C}_3\text{H}_5\text{O}_6 \cdot \text{H}_2\text{O}$	234.05	yelsh viac liq; wh or powd	—	—	—	—	—	—
a311	glycerophosphate, pentahydrate	$\text{Na}_2\text{C}_3\text{H}_5\text{O}_6 \cdot 5\text{H}_2\text{O}$	315.12	wh pl, ac or powd	—	>130	—	—	—	—
a312	gold sulfide	$\text{NaAuS}_2 \cdot 4\text{H}_2\text{O}$	324.08	col, monocl, 1.309	—	—	—	—	—	—
a313	hydride	$\text{NaH}$	24.00	silver need, 1470	0.92	800	—	—	—	—
a314	hydroxide	$\text{NaOH}$	40.00	wh, deliq, 1.3576	2.130	318	1390	—	—	—
a315	iodate	$\text{NaIO}_3$	197.89	wh, rhomb, 4.277	—	—	—	—	—	—
a316	metaperiodate	$\text{NaIO}_4$	213.89	col, tet, 4.174	—	300	—	—	—	—
a317	metaperiodate, trihydrate	$\text{NaIO}_4 \cdot 3\text{H}_2\text{O}$	267.94	col, rhomb; eff	3.219	175	—	—	—	—
a318	paraperiodate	$\text{Na}_2\text{IO}_6$	337.85	wh, rhomb, 1.309	—	800	—	—	—	—
a319	(tri-)peraperiodate	$\text{Na}_3\text{H}_2\text{IO}_7$	293.85	col, hexag, 1.309	—	—	—	—	—	—
a320	iodide	$\text{NaI}$	149.89	col, cub, 1.7745	3.667	661	1304	—	—	—
a321	iodide, dihydrate	$\text{NaI} \cdot 2\text{H}_2\text{O}$	185.92	col, monocl, 1.7745	2.448	752	—	—	—	—
a322	iodoplatinate	$\text{Na}_2\text{PtI}_6 \cdot 6\text{H}_2\text{O}$	1110.59	brn, monocl, 1.7745	3.707	—	—	—	—	—
a323	iridium chloride	Sodium hexachloroiridate, $\text{Na}_3\text{IrCl}_6 \cdot 12\text{H}_2\text{O}$	690.07	olive cr, rhomb or trig-rhomb	—	—	—	—	—	—
a324	iron (III) nitropentacyanide	$\text{Na}_3[\text{Fe}(\text{CN})_5\text{NO}] \cdot 2\text{H}_2\text{O}$	297.95	ruby red, rhomb, 1.605, 1.575, 1.56	1.687	—	—	—	—	—
a325	iron (III) oxalate	$\text{Na}_2[\text{Fe}(\text{C}_2\text{O}_4)_3] \cdot 5\text{H}_2\text{O}$	487.96	grn, monocl, 1.973	—	—	—	—	—	—
a326	iron (III) sulfate	$3\text{Na}_2\text{SO}_4 \cdot \text{Fe}_2(\text{SO}_4)_3 \cdot 6\text{H}_2\text{O}$	934.09	wh, trig, 1.558, 1.613	2.558	—	—	—	—	—
a327	lactate	$\text{NaC}_3\text{H}_5\text{O}_3$	112.06	col or yelsh liq, very hygro	—	17	—	—	—	—
a328	lithium sulfate	$\text{NaLi}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	378.12	col, ditrig, 2.009	—	677 CO <sub>2</sub> 1240/atm	—	—	—	—
a329	magnesium carbonate	$\text{Na}_2\text{CO}_3 \cdot \text{MgCO}_3$	190.31	wh, rhomb, 2.729	—	—	—	—	—	—
a330	magnesium sulfate	Nat. bloedite, $\text{Na}_2\text{SO}_4 \cdot \text{MgSO}_4 \cdot 4\text{H}_2\text{O}$	334.48	col, monocl, 1.488, 1.489	2.23	—	—	—	—	—
a331	magnesium tartrate	$\text{Na}_2\text{Mg}(\text{C}_4\text{H}_4\text{O}_6)_2 \cdot 10\text{H}_2\text{O}$	546.59	wh, monocl pr or powd, 1.488, 1.489	—	—	—	—	—	—